A variety of characteristic derivatives have been made. Thionurate of ammonia $\left[\mathrm{C}_{3} \mathrm{H}_{4} \mathrm{~N}_{3} \mathrm{SO}_{6}\left(\mathrm{NH}_{4}\right)_{2}\right]$ is prepared by acting upon the nitrous derivative with sulphate of ammonia; and amidomalonylurea $\left(\mathrm{C}_{4} \mathrm{H}_{3} \mathrm{NH}_{2} \mathrm{~N}_{2} \mathrm{O}_{3}\right)$, by reducing the former by means of chloride of tin. Alloxantine is inade by treating dibromomalonylurea ( $\mathrm{C}_{4} \mathrm{H}_{2} \mathrm{Br}_{2} \mathrm{~N}_{8} \mathrm{O}_{3}$ ) with hydric sulphide. This can be changed into alloxan ( $\mathrm{C}_{4} \mathrm{H}_{2} \mathrm{~N}_{8} \mathrm{O}_{4}$ ) by adding twice its weight of water and leating to solution with a few drops of nitric acid. Murexide is obtained by treating amidomalonylurea with red oxide of mercury. If tartronic acid ( $\mathrm{C}_{4} \mathrm{H}_{4} \mathrm{O}_{5}$ ) be treated with urea and oxichloride of phosphorus, oximalonylurea is probably obtained, as the product gives a characteristic color by sucressive treatment with nitric acid and ammonia.

On Amalgams of Chromium, Iron Cobalt, Nickel and Manganest, and on a New Process for the Preparation of Metallic Chromium, H. Moissan.-Chromium amalgam is made by acting upon sodium amalgam by solution of protochloride of chromium. This amalgam, heated in a current of hydrogen, gives metallic chromium. An amalgam of manganese is made by decomposing a solution of protochloride of manganese, in the presence of a negative electrode of mercury, by means of a battery. Amalgams of iron, cobalt and nickel, can also be prepared, and from them the pure metals can be reduced.

Aurlysis of some Metallic Fragments taken from Peruvian sepulchres at Apcon, near Lima, A. Terreil.-These analyses were undertaken with the idea that some light might be thrown $u_{1}$ on the condition of metallurgy in this country in the sixteenth century. It will be noticed that the first sample contained chlorine. There was much sea sand in the locality where this metal was found. The analyses are as follows:

| Silver | $\begin{gathered} \text { No. } 1 \\ 77.04 \end{gathered}$ | $\begin{aligned} & \text { No. } 2 . \\ & 33.35 \end{aligned}$ | No. 3. $17.27$ | $\text { No. } 4 .$ | No. 5. trace. |
| :---: | :---: | :---: | :---: | :---: | :---: |
| Gold | trace. | 5.42 | - | - | - |
| Copper | 7.06 | 60.83 | 79.03 | 65.90 | 94.35 |
| Zine | - | - | - | 32.04 | - |
| Iron | - | - | - | 1.05 | - |
| Chlorine. | 15.71 | 0.22 | 2.31 | trace. | trace. |
| Oxygen, sulphur, arsenic, carbonic acid, etc., undet. | 0.19 | 0.18 | 1.39 | 1.01 | 5.53 |
| Quartz sand | - | - | - | - | 0.12 |
|  | 100.00 | 100.00 | 100.00 | 100.00 | 100.00 |

